

**Table VIII**  
**Solubility of Ammonia in Aqueous Salt Solution at 25C\***

Salt	Chemical Formula	Gram Moles NH <sub>3</sub> Soluble in 1 Liter of		
		0.5 Normal Solution	1.0 Normal Solution	1.5 Normal Solution
Potassium Chloride	KCl	0.930	0.866	0.809
Potassium Bromide	KBr	0.950	0.904	0.857
Potassium Iodide	KI	0.970	0.942	0.900
Potassium Hydroxide	KOH	0.852	0.716	0.607
Sodium Chloride	NaCl	0.938	0.889	0.843
Sodium Bromide	NaBr	0.965	0.916	0.890
Sodium Iodide	NaI	0.995	0.992	0.985
Sodium Hydroxide	NaOH	0.876	0.789	0.716
Lithium Chloride	LiCl	0.980	1.008	1.045
Lithium Bromide	LiBr	1.001	1.040	1.090
Lithium Iodide	LiI	1.030	1.094	1.190
Lithium Hydroxide	LiOH	0.865	0.808	0.786
Potassium Fluoride	KF	0.839	0.722	0.626
Potassium Nitrate	KNO <sub>3</sub>	0.923	0.862	0.804
Potassium Nitrite	KNO <sub>2</sub>	0.920	0.855	0.798
Potassium Cyanide	KCN	0.926	0.858	0.802
Potassium Thiocyanate	KSCN	0.932	0.868	0.814
Potassium Sulfate	K <sub>2</sub> SO <sub>4</sub>	0.875	0.772	0.678
Potassium Sulfite	K <sub>2</sub> SO <sub>3</sub>	0.865	0.768	0.675
Potassium Carbonate	K <sub>2</sub> CO <sub>3</sub>	0.788	0.650	0.554
Potassium Oxalate	K <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	0.866	0.771	0.675
Potassium Chromate	K <sub>2</sub> CrO <sub>4</sub>	0.866	0.771	0.675
Potassium Acetate	KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	0.866	0.765	0.685
Potassium Formate	KCHO <sub>2</sub>	0.868	0.760	0.678
Potassium Borate, Meta	KBO <sub>2</sub>	0.814	0.677	0.560
Potassium Acid Phosphate	K <sub>2</sub> HPO <sub>4</sub>	0.860	0.749	0.664
Sodium Sulphide	Na <sub>2</sub> S	0.887	0.795	0.726
+Potassium Chlorate	KClO <sub>3</sub>	0.927		
+Potassium Bromate	KBrO <sub>3</sub>	0.940		
+Potassium Iodate	KIO <sub>3</sub>	0.951		

+ Concentration of these salts is 0.25 Normal.

\* Data of Abegg; Riesenfeld, Z Phys. Ch., 40, 100 (1902)